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Pav

Period 5

Characteristics of Sodium

*General Characteristics*

* Sodium Symbol = NA
* The atomic # is 11
* 22.9898 is the atomic mass of sodium
* Sodium has 11 protons and almost 12 electrons
* NA-23 is the most common of sodium’s isotopes
* Sodium has 12 neutrons in its most common isotope

*Physical characteristics*

* Sodium is solid at room temperature
* Melting point is 97.8 degrees Celsius
* Boiling point is 883 degrees Celsius
* Sodium’s density is 968 kg/m
* Sodium is the 6th most abundant element on earth and is the most abundant of the alkali metals. 2.6% of the Earth’s crust.

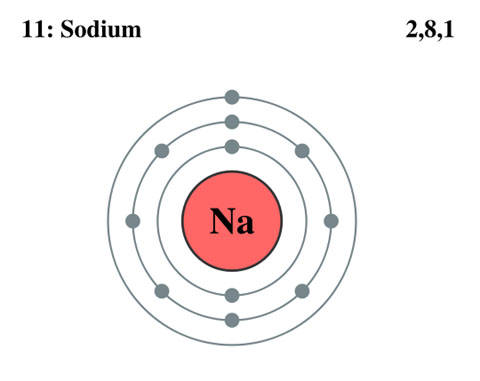
*Background*

* Symbol of sodium comes from Latin “natrium” and the Arabic word “natrun”
* The scientist who discovered the element of sodium is Sir Humphrey Davy
* He discovered it in the year 1807
* Its common uses are household table salt, and sodium compounds are used in glass, soap, paper, chemical and petroleum industries.
* The chemical family of sodium is Alkali Metals

*Chemical Properties*

* Chlorine reacts with sodium to create table salt
* Another element that reacts with sodium is oxygen. The sodium oxidizes in a very quick manner once introduced to oxygen.
* Two compounds that retain sodium are Sodium Chloride (salt) as well as Sodium Hydroxide.

*Conclusive Paragraph on Sodium Chloride*

 Sodium Chloride is very important for animal/human nutrition. Sodium itself is important for retaining fluid balance in the body’s cells. Such as making sure that the cells do not get too flooded with fluid. Sodium is also important for nerve function in the body. Without any sodium, the nerves wouldn’t work the way they’re supposed to. Another use of Sodium and it’s compounds is food preservation, and they are also used as nuclear reactors.

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| <http://pt.chemicalstore.com/Na%20-%20Sodium.html>. <http://chemistry.about.com/od/elementfacts/a/sodium.htm> |  |